



TECHNICAL UNIVERSITY OF MOMBASA

FACULTY OF APPLIED & HEALTH SCIENCES

MATHEMATICS & PHYSICS DEPARTMENT

UNIVERSITY EXAMINATION FOR:

BACHELOR OF TECHNOLOGY IN ENVIRONMENTAL PHYSICS &

RENEWABLE ENERGY

APS 4303: THERMAL PHYSICS II

END OF SEMESTER EXAMINATION

SERIES: MAY 2016

TIME: 2 HOURS

DATE: MAY 2016

Instructions to Candidates

You should have the following for this examination

-Answer Booklet, examination pass and student ID

This paper consists of 4 questions.

Do not write on the question paper. Answer question ONE (compulsory) and any other two questions.

SECTION A (30POINTS)

QUESTION 1

(a) Explain the following terms

(i) Ensemble average

[3points]

(ii) Partition function

[3points]

(iii) Entropy

[3points]

(b) (i) Consider a one-particle system of two states, one of energy 0 and one of energy v . The particles are in thermal equilibrium with a reservoir at temperature \dagger . Compute the energy and heat capacity of the system as a function of then temperature \dagger . [7points]

(ii) If we shift the zero energy and take the energies of the two states as $-\frac{1}{2}v$ and $+\frac{1}{2}v$, compute the partition function and heat capacity of the system and find what the heat capacity looks like in conventional temperature system. [7points]

(c) Consider a model system with N_{\uparrow} spins up and N_{\downarrow} spins down. Let $N = N_{\uparrow} + N_{\downarrow}$; the spin excess is $2s = N_{\uparrow} - N_{\downarrow}$. The entropy is given by (in Stirling's approximation)

$$\dagger(s) \approx -\left(\frac{1}{2}N + s\right) \log\left(\frac{1}{2} + \frac{s}{N}\right) - \left(\frac{1}{2}N - s\right) \log\left(\frac{1}{2} - \frac{s}{N}\right).$$

In a magnetic field B , what would be the free energy and what is the expression for the minimum energy? [7points]

SECTION B

QUESTION 2

(a) For a particle in a box the energy is given by

$$v_n = \frac{\hbar^2}{2m} \left(\frac{f}{L} \right)^2 (n_x^2 + n_y^2 + n_z^2), \text{ where the letters have their usual meanings.}$$

(i) Give the expression for the partition function for this system. [4points]

(ii) What is the expression for the partition function if the spacing between adjacent energy is small in comparison with \dagger . Use the formula,

$$\left(\int_0^\infty dn_x \exp(-r^2 n_x^2) \right)^3 = \frac{f^{3/2}}{8r^3} \quad [6points]$$

(b) The partition function of an ideal of N identical particles is given by

$$Z_N = \frac{1}{N!} (n_Q V)^N, \text{ where } n_Q = \left(\frac{M\dagger}{2f\hbar^2} \right) \text{ and the letters have their usual}$$

meanings.

(i) Determine energy of the gas. [3points]

(ii) Determine the pressure of the gas. [3points]

liii) Determine the entropy of the system. [4points]

QUESTION 3

(a) Give a brief description of the Debye model of heat capacity. [10points]

(b) In the Debye model of heat capacity the total energy is given by

$$U = \int_0^{n_D} dn \frac{\hbar\check{S}}{\exp(\hbar\check{S}/\dagger) - 1}.$$

(i) Determine the total energy in the low temperature limit. [5points]

(ii) Determine the heat capacity C_V in the low temperature regime. [5points]

QUESTION 4

(a) Define the chemical potential of an ideal gas. [4points]

(b) The free energy of a monatomic gas is given by

$$F = -kT \left[\log Z_1 - \log N \right] \text{ where } Z_1 = n_Q V = \left(\frac{M kT}{2\pi \hbar^2} \right)^{3/2} V$$

From this expression determine the chemical potential. [6points]

(c) The differential of entropy is given by $dS(U, V) = \left(\frac{\partial S}{\partial U} \right)_V dU + \left(\frac{\partial S}{\partial V} \right)_U dV$.

If denote the independent values of dU by $(dU)_n$ and dV by $(dV)_n$ the entropy change will be zero.

(i) Determine the expression for the pressure in terms of T, V with U kept constant. [6points]

(ii) From the expression for dS obtain the thermodynamic identities. [4points]

